

Writing Chemical Formulas

1. Write the symbol of the positively charged ion/element first.
2. Write the symbol of the negatively charged ion/element next.
3. Write the oxidation numbers as superscripts behind the symbols.
4. If the sum of the oxidation numbers = 0, skip to step 7.
5. If the sum of the oxidation numbers $\neq 0$, use the criss-cross method to find the correct subscripts so the charge of the compound is 0.
6. Make sure the sum of the oxidation numbers = 0.
7. Write the final formula with symbols and subscripts only!!

Writing Formula Names

1. Write the name of the first element or polyatomic ion.
2. IF a binary compound, write the root of the second element, add *-ide*.
3. IF a polyatomic ion, write the name of the ion unchanged.
4. Check to see if the first element is a transition metal.
5. IF a transition metal, find out what the oxidation number is.
6. IF the transition metal is variable (more than one oxidation number) write the oxidation number as a roman numeral after the name in parentheses.